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## **Quantification of Elements and Compounds**

1. Find the relative molecular mass and molar mass of the following compounds.

i.	CO <sub>2</sub> (Carbon dioxide)	ii.	NaCI (Sodium chloride)		
iii.	CaCO <sub>3</sub> (Calcium carbonate)	iv.	NH₄CI (Ammonium chloride)		
v.	Mg <sub>3</sub> N <sub>2</sub> (Magnesium ni-tride)	vi.	H <sub>2</sub> S (Hydrogen sulphide)		
vii.	AICI <sub>3</sub> (Aluminium chloride)	viii.	(NH <sub>4</sub> ) <sub>2</sub> CO <sub>3</sub> (Ammonium carbonate)		
ix.	CuSO <sub>4</sub> (Copper sulphate)	х.	Na <sub>2</sub> C <sub>2</sub> O <sub>4</sub> (Sodium oxalate)		
xi.	CH₃OH (Methyl alchohol/Methanol)	xii.	CS2 (carbon disulphide)		
xiii.	C <sub>8</sub> H <sub>18</sub> (Octane)	xiv.	CH3COOH (Acetic acid		
xv.	C <sub>12</sub> H <sub>22</sub> O <sub>11</sub> (Sucrose)	xvi.	CO(NH <sub>2</sub> ) <sub>2</sub> (Urea)		
xvii.	C <sub>9</sub> H <sub>8</sub> O <sub>4</sub> (Asprine)	xviii.	HNO <sub>3</sub> (Nitric acid)		
xix.	CCl <sub>4</sub> (Carbon tetrachloride)	xx.	C <sub>8</sub> H <sub>9</sub> NO <sub>2</sub> (Paracetamol)		
(R.A.N	(R.A.M. : H=1. C=12. N=14. 0=16. Na=23. Mg=24. Al=27.S=32.Cl=35.5. S=32.)				

## 2. Calculate the following, considering 67g of Sodium oxalate.

- a. The number of Sodium Oxalate moles
- b. The number of Oxygen moles
- c. Total number of atoms in moles
- d. The number of carbon atoms
- 3. Calculate the following, considering 2.5mol of Ammonium carbonate.
  - e. The mass
  - f. Number of Hydrogen atoms
  - g. Total number of atoms in moles

4. 500mg of Paracetamol is included in the tablet of Paracetamol that use as a pain killer. Find the number of moles of Paracetamol includes in a tablet. Then find the number of moles of each element contains in it.