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Acids, Bases and Salts

• Prepared a list of chemicals that we use in our day to day life.

• Group them as acids, bases and salts/neutrals

Acids	Bases	<u>Salts</u>

Generally salts show neutral properties.

<u>Acids</u>

• Strong acids that are used in laboratory.

H₂SO₄- Sulphuric acid

HNO₃ – Nitric acid

HCI- Hydrochloric acid

Acid is a

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 $HCl (aq) \longrightarrow H^+(aq) + Cl^-(aq)$

• Acids are grouped according to their ability to release H+ ions in an aqueous medium.

Strong acids

Release H+ ions while completely getting ionised in an aqueous medium

Hydrochloric acid Sulphuric acid Nitric acid (HCI) - HCI_(aq) \rightarrow H_{+(aq)} + CI_{-(aq)}

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Week acids

Release H+ ions while getting partially ionised in an aqueous medium

Write the examples for weak acids

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Physical and chemical properties of acids.

- Have corrosion properties.
- Sour in taste.
- Most metals react with acids and release Hydrogen gas.

 $Mg(s) + 2HCl(aq) \longrightarrow MgCl_2(aq) + H_2(g)$

React with carbonates and bicarbonates and release Carbon dioxide gas

 $CaCO_3(s) + 2HCl(aq) \longrightarrow CaCl_2(aq) + H_2O(l) + CO_2(g)$

• Reacts with bases and form neutral salts.

 $H_2SO_4(aq) + 2NaOH(aq) \longrightarrow Na_2SO_4(aq) + 2H_2O(l)$

Write uses of following acids

Sulphuric acid.	Hydrochloric acid	Nitric acid