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Write the answers for the following questions after reading the chapter "**Chemical Bonds**" in your text book.

(1) Complete the following table.

Molecule	Lewis dot structure
H <sub>2</sub>	
Cl <sub>2</sub>	
NH <sub>3</sub>	
H <sub>2</sub> O	

- (2) What the special properties of water due to the intermolecular forces of water molecule?
- (3) The set up relevant to the activity 03 in page number 180 of your text book can be made at home. Use a 9 V battery for dry cell here. Take small amount (30 g) of copper sulphate. Use a jam bottle instead of beaker. Wash your hands after the activity.

Solution	Bulb lighten / not lighten	Electricity conducts/not	Bond type of the compound
Sugar solution			
Copper sulphate solution			
Salt solution			

- 4) Mention four differences between covalent compounds and ionic compounds.
- (5) What are lone pairs in a molecule?
- (6) Draw the structures of the following molecules and mark the lone pairs on them.
  - (i) CO<sub>2</sub>
  - (ii) N<sub>2</sub>
  - (iii) Cl<sub>2</sub>